**CHEMISTRY WORKSHEET** Page 1

|  |  |
| --- | --- |
| **NAME THE FOLLOWING** | **WRITE FORMULAS FOR THE FOLLOWING** |
| 1. Li2O | 1. zinc hydroxide |
| 2. KMnO4 | 2. gold (I) acetate |
| 3. (NH4)2CO3 | 3. diselenium pentoxide |
| 4. Fe(OH)3 | 4. carbon monoxide |
| 5. N2O3 | 5. nickel (III) sulfite |
| 6. Cu(NO3)2 | 6. cobalt (II) nitrate |
| 7. SeO2 | 7. strontium chloride |
| 8. CCl4 | 8. silicon tetrachloride |
| 9. BaCr2O7 | 9. lead (II) carbonate |
| 10. SrCrO4 | 10. copper (II) sulfate |
| 11. P2O5 | 11. tin (IV) dichromate |
| 12. Mn3(PO4)7 | 12. phosphorus triiodide |
| 13. Ag3PO3 | 13. iron (III) nitrite |
| 14. CO2 | 14. potassium permanganate |
| 15. AlCl3 | 15. sodium chlorate |
| 16. NaOH | 16. diarsenic hexaoxide |
| 17. Sn(CN)4 | 17. tungsten (V) hypochlorite |
| 18. Zn(NO3)2 | 18. barium chromate |
| 19. SO3 | 19. potassium hydroxide |
| 20. NaC2H3O2 | 20. ruthenium (III) sulfite |

**Write the formula for each of the acids listed below:** Page 2

|  |  |
| --- | --- |
| **NAME THE FOLLOWING** | **WRITE FORMULAS FOR THE FOLLOWING** |
| 1. HClO4 | 1. Nitric acid |
| 1. H3PO4 | 1. Chloric acid |
| 1. HCl (aq) | 1. Acetic acid |
| 1. H2SO4 | 1. Hydrobromic acid |
| 1. HNO2 | 1. Sulfurous acid |
| 1. HI (aq) | 1. Chlorous acid |
| 1. HC2H3O2 | 1. Hydrochloric acid |
| 1. HF (aq) | 1. Phosphoric acid |
| 1. H3PO3 | 1. Nitrous acid |
| 1. HClO3 | 1. Hydrofluoric acid |
| 1. H2CO3 | 1. Perchloric acid |
| 1. H2SO3 | 1. Hydroiodic acid |
| 1. HClO2 | 1. Phosphorous acid |
| 1. HNO3 | 1. Carbonic acid |
| 1. HBr (aq) | 1. Sulfuric acid |

**Chemistry: *Errors in Chemical Formulas and Nomenclature*** Page 3

*Each of the following formulas or chemical names contains an error. Correct each example.*

1. aluminum (III) iodide 1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

2. Al3O2 2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

3. (NH4)3Cl2 3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

4. lead (III) oxide 4. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

5. leadic oxide 5. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

6. stannous (II) bromide 6. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

7. monocalcium dioxide 7. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

8. K(ClO2) 8. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

9. (OH)3Al 9. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

10. Cr4(CN)3 10. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

11. Pb(NO3)3 11. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

12. copper (I) chloric 12. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

13. ferric (III) oxide 13. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

14. NiNO32 14. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

15. Mg2F 15. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

16. ironic oxide 16. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

17. ClK 17. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

18. Ca2O2 18. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

19. sodium (I) fluoride 19. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

20. dititanium tetroxide 20. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

21. dihydrogen oxide 21. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

22. magnesium dihydroxide 22. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

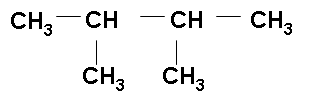
23. cuprous (I) sulfate 23. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

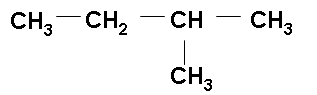
24. strontium difluoride 24. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

25. NO3Na 25. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

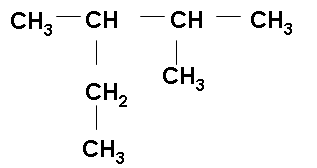
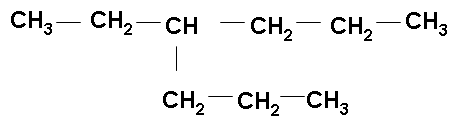
**Name the following alkanes** Page 4

1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

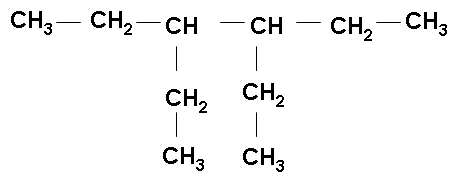
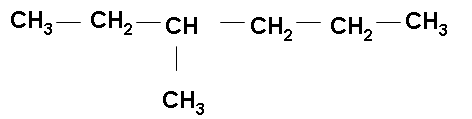




3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 4. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_



5. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ 6. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_



**Draw these molecules:**

7. 2, 2, 4-trimethyl pentane 8. 2, 3-diethyl 4-propyl octane

9. 3, 4-diethyl 2,2-dimethyl 3-propyl decane